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Semiempirical Study of Rare Gas and Rare Gas-Hydrogen Ionic Clusters:
 R_n^+ , $(R_n H)^+$ and $(R_n H_2)^+$ for $R = Ar, Xe$

by

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Semiempirical Study of Rare Gas and Rare Gas-Hydrogen Ionic

Clusters: R_n^+ , $(R_nH)^+$ and $(R_nH_2)^+$ for $R = Ar, Xe$

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Abstract

The ionic rare gas clusters Ar_n^+ and Xe_n^+ and rare gas-hydrogen clusters $(Ar_nH)^+$, $(Ar_nH_2)^+$, $(Xe_nH)^+$ and $(Xe_nH_2)^+$ are studied by the semiempirical diatomics-in-ionic-systems (DIIS) method. The Ar_n^+ clusters ($n > 3$) are seen to have a structure of a linear Ar_3^+ core surrounded by $n-3$ neutral or almost neutral Ar atoms. For Xe_n^+ ($n > 3$), a symmetrical Xe_4^+ ionic core with the geometry of regular pyramid is formed. The rare gas-hydrogen clusters with one H atom have a simple $R_k(RH)^+$ structure with k neutral rare gas atoms attracted to the $(RH)^+$ molecule by polarization forces. Two H atoms can bind with Ar atoms to form quasistable clusters $Ar_nH_2^+$ which dissociate to $(n-1)Ar + H + (ArH)^+$ through a high barrier of roughly 0.75 eV. Two H atoms and one Xe^+ ion are shown to form a collinear valence-bound $(XeHH)^+$ cluster whose dissociation energy is 0.46 eV.



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I. Introduction

The simplest small clusters are formed by (i) rare gas atoms,¹ (ii) rare gas atoms and nonpolar molecules²⁻⁴ and (iii) rare gas atoms and polar molecules.⁵⁻⁷ In all these clusters, electrons are strongly localized on the component molecules and rare gas atoms such that the particles are bound together only by dispersion (van der Waals) and in case (iii), weak polarization forces. Such clusters are usually called van der Waals (VDW) clusters. The rare gas atoms can also form, however, clusters in which electrons are partly delocalized, contributing to some valence forces between the particles involved. In particular, such delocalization takes place in neutral rare gas-halogen and ionic rare gas clusters.

As an example of neutral rare gas-halogen clusters, we can mention Xe_nCl . The ground-state Xe_nCl clusters usually seem to be of the weakly-bound VDW-type. However, ab initio calculations of the XeCl complex⁸ show that the large electron affinity of the Cl atom leads to a small but noticeable electron delocalization which contributes about 50% of the XeCl dissociation energy.⁹ This delocalization affects the structure of the Xe_nCl clusters and is responsible for the coupling between their ground state and excited ionic (i.e., charge transfer, Xe_n^+Cl^-) states.⁹ In these excited states the strongly-bound quasistable molecules Xe_2^+Cl^- are formed which have been studied experimentally both in the gaseous¹⁰⁻¹² and solid¹³ phases. The Xe_2^+Cl^- molecule can attract more Xe atoms, mainly within Xe solid, forming Xe_n^+Cl^- ($n > 2$) clusters.^{9,14-15}

The ionic rare gas clusters R_n^+ have been studied extensively both experimentally and theoretically. The smallest clusters R_3^+ have been detected, in particular, in Ar,¹⁶⁻²¹ Xe²² and Kr.²³ Ab initio calculations of Ne_3^+ and Ar_3^+ ,²⁴⁻²⁵ as well as semiempirical calculations of Ar_3^+ ²⁶⁻²⁸ and

Xe_3^+ ,^{27,29} show that the triatomic clusters R_3^+ have a symmetrical linear structure with the central atom bearing roughly 50% of the charge. Clusters with more than three atoms are found for all rare gas atoms, i.e., He_n^+ ,³⁰ Ne_n^+ ,³¹ Ar_n^+ ,^{27,32-36} Kr_n^+ ,³⁷ and Xe_n^+ .^{38,39} The large rare gas clusters R_n^+ with $n > 10$ demonstrate some magic numbers in their stabilization energy dependence on n .^{33,35,37-43} The structure of the R_n^+ clusters with $n > 3$ is not completely clear. According to an experimental study,³⁵ the Ar_3^+ ion is the core in the Ar_n^+ clusters. Theoretical investigations which take into account the electron delocalization support this finding, at least for Ar_n^+ ²⁶ and Xe_n^+ ,²⁹ and show that R_4^+ ions can also form a core of R_n^+ ($n > 4$) clusters. The involvement of R_3^+ and R_4^+ ions in the R_n^+ cluster formation allows one to conclude that models with the charge localized on one³³ or two^{41,42} centers are unsatisfactory, at least for small R_n^+ clusters. As shown in Ref. 27 and 29, the charge can be shared even by more than four atoms.

Rare gas atoms can form not only the homonuclear clusters R_n^+ considered above but also ionic heteronuclear clusters which consist of either different rare gas atoms or rare gas atoms and valence-active atoms. The charge distribution and the structure of heteroatomic ionic clusters depend strongly on the relation between the ionization potentials of the atoms involved. For example, the $(\text{Xe}_n \text{Ne}_m)^+$ cluster is obviously formed by a Xe_n^+ ionic core and neutral Ne atoms, since the ionization potential of Ne is much higher than that of Xe. In the case of valence-active atoms such as hydrogen, the situation becomes more complicated because of the formation of the valence-bonded molecules H_2 , H_2^+ , H_3^+ and $(\text{RH})^+$. The $(\text{RH}_2)^+$ systems have been studied in the context of the $\text{R}^+ + \text{H}_2$ and $\text{R} + \text{H}_2^+$ chemical reactions (see, for example, Refs. 44-46). The $(\text{ArH}_2)^+$ potential energy surface has been calculated by the DIM method.⁴⁷ The study of the rare gas-hydrogen ionic clusters has

concentrated mainly on ArH_3^+ clusters,⁴⁸ whose structure, by analogy with HeH_3^+ ,⁴⁹ has been suggested to be a H_3^+ triangle with Ar at the vertex. This structure is supported by ab initio calculations.⁵⁰ The ionization potential of Ar (15.76 eV) is much higher than that of H (13.6 eV), so that the charge is located on the H atoms. A similar charge distribution is expected for the $(\text{H}_2)_n\text{H}_3^+$ ionic clusters, which have been studied experimentally,⁵¹ since the ionization potential of H_2 (15.43 eV) is also much higher than the H ionization potential. Ab initio calculations of rare gas-hydrogen clusters with one H atom have been performed for Ne_nH^+ clusters.⁵² Due to the large difference between the Ne and H ionization potentials, the charge is strongly localized on H^+ , and the Ne atoms are bound to H^+ mainly by polarization forces. A similar situation also occurs for clusters formed by aromatic cations and rare gas atoms. Since the ionization potentials of aromatic molecules are much smaller than those of rare gas atoms, the positive charge is located on the aromatic molecules, which thus attracts the rare gas atoms by polarization forces.⁵³⁻⁵⁵

The brief survey presented above shows that the study of ionic clusters is a rapidly-developing field of research. Some of the experimental results, especially the discovery of the new (cluster) class of reactions,^{56,57} suggest possible practical applications of ionic clusters, such as in catalysis.

In this study we shall present a quantum chemical treatment of the relatively simple ionic clusters R_n^+ , $(\text{R}_n\text{H})^+$ and $(\text{R}_n\text{H}_2)^+$. These systems are bound together not only by electrostatic forces but sometimes also by valence forces resulting from the electron (charge) delocalization, as discussed above for R_3^+ ions. In order to study theoretically the systems with ions where both electrostatic polarization and charge delocalization are of importance, we have previously developed the semiempirical diatomics-in-ionic-systems (DIIS)

method.⁹ This method was first applied to rare gas-halogen systems with excited ionic states like Xe_n^+Cl^- ^{9,15} and $\text{Xe}_n^+\text{HCl}^-$,⁷ and we shall now use it to calculate the electronic structure of rare gas and rare gas-hydrogen ionic clusters.

In the next section we shall describe briefly the DIIS method, as well as some modifications which one needs to apply it to the systems under consideration. Sections III and IV deal with calculations of Ar_n^+ , $(\text{Ar}_n\text{H})^+$ and $(\text{Ar}_n\text{H}_2)^+$ and of Xe_n^+ , $(\text{Xe}_n\text{H})^+$ and $(\text{Xe}_n\text{H}_2)^+$, respectively. Section V gives the Conclusions, and the Appendix presents the diatomic potentials.

II. Diatomics-in-Ionic-Systems (DIIS) Method

A. Closed-Shell Atoms with Electron Deficiency

Let us consider a polyatomic system $(\text{A}_1\text{A}_2\ldots\text{A}_J)^+$ consisting of J closed-shell atoms or ions A_i with one electron deficiency (hole) delocalized, generally speaking, between these atoms (ions). For example, the neutral rare gas-halogen system R_nX is considered in DIIS as $(\text{R}_n\text{X}^-)^+$. The spin of a system with one electron deficiency is $S = \frac{1}{2}$. The DIIS wave function is presented as a linear combination of diabatic polyatomic wave functions with localized electron deficiency as⁹

$$\Phi = \sum_{i=1}^J \sum_{m=1}^{M_i} C_{im} \Phi_{im} \quad (1)$$

where Φ_{im} is the diabatic wave function of the configuration $\text{A}_1\text{A}_2\ldots\text{A}_{im}^+\ldots\text{A}_J$, with the index m indicating the orientation of the A_i^+ open shell ($M_i = 3$ for a P-symmetry A_i^+ shell). For example, R_2X is described by nine diabatic

configurations, namely R^+RX^- , RR^+X^- and RRX for $m = 1, 2, 3$. The rare gas ionic cluster R_n^+ is described by $3n$ diabatic configurations $R...R^+...R$. An ionic system consisting of n rare gas atoms and one hydrogen atom $(R_nH)^+$ can be also treated by the DIIS method if the electron affinity of H is neglected and singlet states only are considered.⁷ The $(R_nH)^+$ system is described by $3n + 1$ diabatic configurations, namely $R...R^+...RH$ ($m = 1, 2, 3$) and $R...R...RH^+$ ($m = 1$).

The diagonal matrix elements of the DIIS wave function (1) have the physical interpretation as the energies of the diabatic configurations with fixed localization of the electron deficiency. If we consider only the pairwise interactions between atoms (ions), then the diagonal matrix elements are presented as a sum of diatomic interactions,

$$H_{im,im} = \sum_j W_j + \sum_{j_1} \sum_{j_2(>j_1)} W_{j_1j_2} + U_i + \sum_j U_{im,j}, \quad j, j_1, j_2 \neq i, \quad (2)$$

where W_j is the energy of the closed-shell atom A_j , $W_{j_1j_2}$ is the diatomic potential between two closed-shell atoms, $A_{j_1} - A_{j_2}$, U_i is the energy of atom A_i^+ with an electron deficiency, and $U_{im,j}$ is the diabatic potential between the closed-shell atom A_j and the atom with electron deficiency, A_i^+ . The transformation of the potentials between m -oriented orbitals to the Σ and Π potentials is described in Ref. 9, where the expressions for the off-diagonal matrix elements are given as well. The atom H is considered here as a closed shell atom. In the case of ionic systems, like R_n^+ or $(R_nH)^+$, the closed-shell atoms are neutral, and the atoms with electron deficiency are the ions R^+ or H^+ . The diatomic potentials between neutral atoms, W_{RR} and W_{RH} , are taken to be known from empirical studies or ab initio calculations. The diabatic

potentials between neutral atoms and ions, such as R^+-R , R^+-H , $R-H^+$, have to be calculated since only adiabatic potentials, such as $W_{R_2}^+$ and $W_{(RH)^+}$, are the physically-defined values.

The diabatic potentials $U_{im,j}$ can be found by considering the diatomic fragments within the DIIS approximation and solving the inverse problem of the 2×2 matrix eigenvalues, i.e., calculating the matrix elements for given eigenvalues. In the case of the homonuclear fragment R_2^+ , the 2×2 matrix of the fragment is

$$\begin{vmatrix} U - E & V \\ V & U - E \end{vmatrix} = 0, \quad (3)$$

where U is the R^+-R diabatic potential and V is the exchange term in the $(R^+R)-(RR^+)$ coupling. Substituting into (3) the known adiabatic potentials $E^{(1)} = {}^2\Sigma_u$ and $E^{(2)} = {}^2\Sigma_g$ or $E^{(1)} = {}^2\Pi_g$ and $E^{(2)} = {}^2\Pi_u$, we can easily find the diabatic R^+R potential U and the exchange term V for Σ - or Π -symmetry. For the case of the heteroatomic fragment $(RH)^+$, there are two different diabatic potentials, R_{RH}^+ and U_{RH}^+ ,

$$\begin{vmatrix} U_{RH}^+ - E & V \\ V & U_{RH}^+ - E \end{vmatrix} = 0. \quad (4)$$

In order to find U_{RH}^+ , U_{RH}^+ and V for two given adiabatic potentials, for example, $E^{(1)} = X^1\Sigma$ and $E^{(2)} = B^1\Sigma$, one needs an extra empirical or ab initio value such as the static or transition dipole moment. After the diabatic

potentials $U_{im,j}$ and the exchange terms $V_{i_1 m_1, i_2 m_2}$ are determined from Eqs. (3) and (4), the diagonal matrix elements (2) and the off-diagonal matrix elements (presented in Ref. 9) can be calculated, and the wave function (1) and its energy spectrum can be found.

In the wave function presentation (1), the spin-orbit coupling is not taken into account, at least directly. Indirectly, the energy level shifts resulting from the spin-orbit coupling are incorporated into the system energy via the empirical or ab initio diatomic potentials which are used in the DIIS calculations as the input. By neglecting the spin-orbit coupling in the wave function, we are losing some of the excited states and are excluding from our consideration effects such as S-P mixing in collinear configurations.

B. Polarization Energy

In applying Eqs. (2)-(4), we need not care about the polarization component of the ion-atom interactions, since it is included in the empirical or ab initio potentials which we are suggesting to use. However, the polarization of atoms gives rise to the electrostatic dipole-dipole interaction which is of three-particle origin (ion-atom-atom) and cannot be considered in a pairwise way. The energy of the dipole-dipole interaction is usually small, although it is not necessarily negligible, and hence we intend to take it into account. Including this energy as a separate term Δ_i within the diagonal matrix element (2) and taking as zero the energy of the separated neutral atoms, we obtain

$$H_{im,im} = I_i + \sum_{j_1} \sum_{j_2 (> j_1)} W_{j_1 j_2} + \sum_j U_{im,j} + \Delta_i, \quad j, j_1, j_2 \neq i, \quad (5)$$

where I_i is the ionization potential of the i -th atom, and

$$\Delta_i = \sum_{j_1} \sum_{j_2(>j_1)} \Delta_{i,j_1,j_2} \quad (6)$$

with Δ_{i,j_1,j_2} as the dipole-dipole interaction between two neutral atoms polarized by the ion with index i .¹⁴ Since the terms Δ_{i,j_1,j_2} depend on the diatomic polarization energies P_{ij} , we must calculate them. P_{ij} are used for the Δ_i calculation only, so that their effect is supposedly weak, which reduces significantly the requirements for their accuracy.

In the dipole approximation the ion-atom polarization energy decreases with the fourth power of the separation R ,

$$P_{ij} = -C_j/R_{ij}^4, \quad C_j = e^2\alpha_j/2 \quad (7)$$

where α_j is the polarizability of the j -th atom. The dipole approximation (7) is justified provided that atoms do not overlap one with another. From physical reasons it is clear that the overlap between an ion and atom has a damping effect on the polarization energy P_{ij} . We suggest taking this account into effect by multiplying the polarizability α_j in Eq. (7) by a damping function $\gamma(R_{ij})$,

$$\bar{\alpha}_j(R_{ij}) = \alpha_j \gamma(R_{ij}) \quad (8)$$

A suggested form for the functional dependence on R_{ij} is

$$\gamma(R_{ij}) = \left[1 + \left(\frac{r_i + r_j}{R_{ij}} \right)^{12} \right]^{-1/3} \quad (9)$$

where r_i and r_j are the ionic and atomic radii, respectively. For the case of the interaction between polarized neutral atoms, both r_i and r_j are atomic radii. When the interatomic separation R_{ij} is larger than the sum of atomic radii (non-overlapping atoms), the function γ is close to unity, i.e., the dipole approximation is valid. When the interatomic separation is smaller than the sum of atomic radii (overlapping atoms), the damping function decreases as R_{ij}^4 , so that the ion-atom polarization energy (7) is saturated with a limiting value of $-C_j/(r_i+r_j)^4$. The suggested form (9) for the R_{ij} dependence is of course arbitrary, but at least it describes in a proper way the asymptotic behavior. If the asymptotic conditions are fulfilled, the form of the dependence of γ on R_{ij} does not imply much about the results. According to our estimations, different reasonable models of the damping function $\gamma(R_{ij})$ vary the energy within a range of 0.03 eV only, mostly even much less. For the values r_i and r_j in (9), we shall use the atomic VDW radii.

C. $(R_n H_2)^+$ System

The ionic rare gas-hydrogen system $(R_n H_2)^+$ is described in the DIIS method as n closed-shell atoms (R) and two open-shell atoms (H) with one positively-charged hole. Each of the two diabatic configurations $R \dots R H^+ H$ with a hole located on one of the H atoms is described by a wave function with a fixed $s = \frac{1}{2}$ spin on the only open-shell atom (H). The situation is much more complicated in the diabatic configuration $R \dots R^+ \dots R H H$ when the hole is located on one of the n rare gas atoms. In this configuration there are three open-shell atoms (R^+, H, H) with spins $s = \frac{1}{2}$, which make two different spin states for a total spin $S = \frac{1}{2}$. Consequently, each $R \dots R^+ \dots R H H$ configuration with fixed hole location is described by six diabatic wave functions which

differ from another not only by p-orbital orientation, like in expansion (1), but also by spin eigenfunctions. Taking into account the hole delocalization between all rare gas and hydrogen atoms, one expresses the DIIS wave function as the superposition of $6n+2$ diabatic wave functions,

$$\Phi = \sum_{i=1}^n \sum_{m=1}^3 \sum_{k=1}^2 C_{imk} \Phi_{imk} + \sum_{j=1}^2 C_{n+j} \Phi_{n+j} \quad (10)$$

where Φ_{n+j} describes the diabatic configuration with the charge on an H atom, $R \dots RH^+H$, Φ_{imk} describes the diabatic configuration $R_{(1)} \dots R_{(i)}^+ \dots R_{(n)} HH$ (the numbers in parentheses are the atomic indices) with the m -orientation of the $R_{(i)}^+$ P-shell, and k indicates the spin configuration of the three $(R_{(i)}^+, H, H)$ uncoupled electrons. Following the usual DIM presentation of a triatomic wave function,^{47,58} we express the diabatic wave functions in terms of one-electron orbitals as

$$\Phi_{i11} = \frac{1}{\sqrt{2}} (|p_{ix}p_{iy}\bar{p}_{iy}p_{iz}\bar{p}_{iz}s_1\bar{s}_2| - |p_{ix}p_{iy}\bar{p}_{iy}p_{iz}\bar{p}_{iz}\bar{s}_1s_2|) \prod_{j(\neq i)}^n \phi_j \quad (11)$$

$$\begin{aligned} \Phi_{i12} = \frac{1}{\sqrt{6}} (2|\bar{p}_{ix}p_{iy}\bar{p}_{iy}p_{iz}\bar{p}_{iz}s_1s_2| - |p_{ix}p_{iy}\bar{p}_{iy}p_{iz}\bar{p}_{iz}s_1\bar{s}_2| \\ - |p_{ix}p_{iy}\bar{p}_{iy}p_{iz}\bar{p}_{iz}\bar{s}_1s_2|) \prod_{j(\neq i)}^n \phi_j \end{aligned} \quad (12)$$

$$\phi_j = |p_{jx}\bar{p}_{jx}p_{jy}\bar{p}_{jy}p_{jz}\bar{p}_{jz}|, \quad j = 1, 2, \dots, n \quad (13)$$

$$\Phi_{n+1} = s_2 \prod_j^n \phi_j, \quad \Phi_{n+2} = s_1 \prod_j^n \phi_j \quad (14)$$

where the bars denote the negative spin orientation, p_{ix} , p_{iy} , p_{iz} are the p-orbitals of the i -th rare gas atom, and s_1 and s_2 are the 1s orbitals of the hydrogen atoms. The index $m = 1$ in ϕ_{i11} and ϕ_{i12} indicates the x-orientation of the $R_{(i)}^+$ shell. The diabatic functions ϕ_{imk} for $m = 2$ (y-orientation) and $m = 3$ (z-orientation) are expressed by equations similar to Eqs. (11)-(12).

The p-orbitals of the diabatic functions (11)-(13) are oriented along fixed x,y,z-axes. When the diabatic functions are used to form the H-matrix, the x,y,z oriented p-orbitals have to be transformed to Σ and Π orientations of diatomic fragments.⁹ This transformation for arbitrary located atoms is performed by an S matrix whose elements are

$$S_{ij,x\Sigma} = (x_j - x_i)/R_{ij} \quad , \quad S_{ij,x\Pi} = r_{ij}/R_{ij} \quad , \quad S_{ij,x\pi} = 0 \quad , \quad (15a)$$

$$S_{ij,y\Sigma} = (y_j - y_i)/R_{ij} \quad , \quad S_{ij,y\Pi} = -(x_j - x_i)(y_j - y_i)/(R_{ij}r_{ij}) \quad ,$$

$$S_{ij,y\pi} = (z_j - z_i)/r_{ij} \quad , \quad (15b)$$

$$S_{ij,z\Sigma} = (z_j - z_i)/R_{ij} \quad , \quad S_{ij,z\Pi} = -(x_j - x_i)(z_j - z_i)/(R_{ij}r_{ij}) \quad ,$$

$$S_{ij,z\pi} = -(y_j - y_i)/r_{ij} \quad , \quad (15c)$$

$$r_{ij} = \sqrt{(y_j - y_i)^2 + (z_j - z_i)^2} \quad , \quad (16)$$

where x_i , y_i and z_i are the i -th atom coordinates, R_{ij} is the interatomic separation, and the Π -orientation of the $R_{(i)}^+$ P-shell is indicated by Π when located in the plane formed by \vec{R}_{ij} and the x-axis, and by π when perpendicular

to this plane. We shall also use later the following combinations of the S values:

$$T_{ij,m} = S_{ij,m\Pi}^2 + S_{ij,m\pi}^2 \quad (17a)$$

$$T_{ij,m_1 m_2} = S_{ij,m_1\Pi} S_{ij,m_2\Pi} + S_{ij,m_1\pi} S_{ij,m_2\pi} \quad (17b)$$

Applying both the DIIS⁹ and DIM^{47,58} approaches and using the diabatic functions (11)-(14), we obtain the $(6n+2) \times (6n+2)$ H-matrix which we shall simply display without going into the details of its construction. The diagonal matrix element for $R_{(1)} \dots R_{(i)}^+ \dots R_{(n)} H_{(n+1)} H_{(n+2)}$ is

$$\begin{aligned} H_{imk,imk} = & I_R + \sum_{j_1=1}^{n-1} \sum_{(j_2 > j_1)}^n W_{j_1 j_2} + \sum_{j=1}^n [S_{ij,m\Sigma}^2 U_{ij,\Sigma} + T_{ij,m} U_{ij,\Pi}] \\ & + \sum_{j_1=1}^n \sum_{j=1}^2 W_{j_1, n+j} + \sum_{j=1}^2 [S_{i, n+j, m\Sigma}^2 O_{i, n+j, \Sigma}^{(k)} + T_{i, n+j, m} O_{i, n+j, \Pi}^{(k)}] \\ & + W_{n+1, n+2}^{(k)}, \quad j, j_1, j_2 \neq i, \quad k = 1, 2, \end{aligned} \quad (18)$$

$$O_{i, n+j, \Sigma}^{(1)} = \frac{1}{4} {}^1U_{i, n+j, \Sigma} + \frac{3}{4} {}^3U_{i, n+j, \Sigma} \quad (19a)$$

$$O_{i, n+j, \Sigma}^{(2)} = \frac{3}{4} {}^1U_{i, n+j, \Sigma} + \frac{1}{4} {}^3U_{i, n+j, \Sigma}, \quad (19b)$$

and the same for $O_{i, n+j, \Pi}^{(k)}$. In Eqs. (18)-(19), I_R is the rare gas ionization potential, $W_{j_1 j_2}$ are the R-R potentials, $W_{j_1, n+j}$ are the R-H potentials,

$W_{n+1,n+1}^{(k)}$ are the H-H ${}^1\Sigma$ ($k = 1$) and ${}^3\Sigma$ ($k = 2$) potentials, $U_{ij,m\Sigma}$ and $U_{ij,m\Pi}$ are the $R_{(i)}^+ - R_{(j)}$ diabatic potentials, and ${}^1U_{i,n+j,\Sigma}$ (${}^1U_{i,n+j,\Pi}$) and ${}^3U_{i,n+j,\Sigma}$ (${}^3U_{i,n+j,\Pi}$) are the Σ -symmetry (Π -symmetry) $R^+ - H$ diabatic potentials in the singlet and triplet states, respectively. The diagonal matrix element for $R \dots R H_{(n+1)}^+ H_{(n+2)}$ is

$$H_{n+1,n+1} = I_H + \sum_{j_1=1}^{n-1} \sum_{j_2(>j_1)}^n W_{j_1 j_2} + \sum_{j=1}^n (W_{j,n+2} + U_{n+1,j}) + U_{n+1,n+2} \quad (20)$$

and the same for $H_{n+3,n+2}$. In Eq. (20), $U_{n+1,j}$ are the $H^+ - R$ diabatic potentials, and $U_{n+1,n+2}$ is the $H_{(n+1)}^+ - H_{(n+2)}$ diabatic potential.

The off-diagonal matrix element between two $R_{(1)} \dots R_{(i)}^+ \dots R_{(n)} H_{(n+1)} H_{(n+2)}$ configurations with different ($k = 1, 2$) spin orientations are:

$$H_{im_1, im_2} = \frac{\sqrt{3}}{4} \sum_{j=1}^2 (-1)^{j-1} [S_{i,n+j,m\Sigma}^2 ({}^3U_{i,n+j,\Sigma} - {}^1U_{i,n+j,\Sigma}) + T_{i,n+j,m} ({}^3U_{i,n+j,\Pi} - {}^1U_{i,n+j,\Pi})] \quad (21)$$

The off-diagonal matrix elements between two

$R_{(1)} \dots R_{(i)}^+ \dots R_{(n)} H_{(n+1)} H_{(n+2)}$ configurations with different (m_1, m_2) orbital orientations are

$$H_{im_1 k, im_2 k} = \sum_{j=1}^n [S_{ij,m_1 \Sigma} S_{ij,m_2 \Sigma} U_{ij,\Sigma} + T_{ij,m_1 m_2} U_{ij,\Pi}]$$

$$+ \sum_{j=1}^2 [S_{i,n+j,m_1\Sigma} S_{i,n+j,m_2\Sigma} O_{ij,\Sigma}^{(k)} + T_{i,n+j,m_1m_2} O_{ij,\Pi}^{(k)}] \quad , \quad j = 1 \quad (22)$$

$$\begin{aligned} H_{im_11,im_22} - H_{im_2,im_21} - \frac{\sqrt{3}}{4} \sum_{j=1}^2 (-1)^{j-1} S_{i,n+j,m_1\Sigma} S_{i,n+j,m_2\Sigma} \\ \times ({}^3U_{i,n+j,\Sigma} - {}^1U_{i,n+j,\Sigma}) + T_{i,n+j,m_1m_2} ({}^3U_{i,n+j,\Pi} \\ - {}^1U_{i,n+j,\Pi}) \quad . \end{aligned} \quad (23)$$

The off-diagonal matrix elements between the configurations

$R_{(1)} \dots R_{(i_1)}^+ \dots R_n H_{(n+1)} H_{(n+2)}$ and $R_{(1)} \dots R_{(i_2)}^+ \dots R_n H_{(n+1)} H_{(n+2)}$ are

$$\begin{aligned} H_{i_1m_11,i_2m_21} - H_{i_1m_12,i_2m_22} - S_{i_1i_2m_1\Sigma} S_{i_2i_1m_2\Sigma} V_{i_1i_2\Sigma} \\ + (S_{i_1i_2m_1\Pi} S_{i_2i_1m_2\Pi} + S_{i_1i_2m_1\pi} S_{i_2i_1m_2\pi}) V_{i_1i_2\pi} \end{aligned} \quad (24)$$

$$H_{i_1m_11,i_2m_22} = 0 \quad , \quad (25)$$

where $V_{i_1i_2\Sigma}$ and $V_{i_1i_2\Pi}$ are the exchange terms of the

$(R_{(i_1)}^+ R_{(i_2)}) - (R_{(i_1)} R_{(i_2)}^+)$ coupling (see Eq. (3)). The off-diagonal matrix elements between the configuration $R_{(1)} \dots R_{(i)}^+ R_{(n)} H_{(n+1)} H_{(n+1)}$ and the configurations $R_{(1)} \dots R_{(i)} \dots R_{(n)} H_{(n+1)}^+ H_{(n+2)}$ and $R_{(1)} \dots R_{(i)} \dots R_{(n)} H_{(n+1)} H_{(n+2)}^+$ are

$$H_{im1,n+j} = -S_{i,n+j,m,\Sigma} V_{i,n+j,\Sigma} \quad , \quad j = 1,2 \quad (26)$$

$$H_{im2,n+j} = \sqrt{3}(-1)^{j-1} H_{im1,n+j}, \quad j = 1, 2 \quad (27)$$

where $V_{i,n+j,\Sigma}$ is the exchange term in the $(R^+H)-(RH^+)$ coupling (see Eq. (4)). The Π -component of the coupling is zero. The off-diagonal matrix element between the configurations $R_{(1)} \dots R_{(n)} H_{(n+1)}^+ H_{(n+1)}$ and $R_{(1)} \dots R_{(n)} H_{(n+1)}^+ H_{(n+2)}^+$ is equal to the exchange term of the $(H^+H)-(HH^+)$ coupling in the H_2^+ molecule (see Eq. (3)),

$$H_{n+1,n+2} = V_{n+1,n+2} \quad (28)$$

When the delocalization of the charge is neglected (i.e., setting the exchange terms V equal to zero), we obtain for $R \dots R^+ \dots RHH$ the usual DIM matrix elements.

The diabatic potentials $U_{ij,\Sigma}$ and $U_{ij,\Pi}$ of the R^+R diatomic fragments (Eqs. (18) and (22)), as well as $U_{n+1,n+2}$ of the H^+H fragment (Eq. (20)), are found by the inverse solution of Eq. (3), as described in Subsection II.A above. The situation is much more complicated for the diabatic potentials ${}^1U_{i,n+j,\Sigma}$, ${}^3U_{i,n+j,\Sigma}$ (Eqs. (19), (21) and (23)) and $U_{n+1,j}$ (Eq. (20)) of the R^+H and RH^+ diatomic fragments. In our approach, the $(RH)^+$ diatomic fragment is described by a 3×3 matrix which is obtained by removing to infinity all other atoms. Using Eqs. (18), (20), (26) and (27), we have

$$\begin{vmatrix} \frac{1}{4} {}^1U_{i,n+1} + \frac{3}{4} {}^3U_{i,n+1} - E & \frac{\sqrt{3}}{4} ({}^3U_{i,n+1} - {}^1U_{i,n+1}) & V_{i,n+1} \\ \frac{\sqrt{3}}{4} ({}^3U_{i,n+1} - {}^1U_{i,n+1}) & \frac{3}{4} {}^1U_{i,n+1} + \frac{1}{4} {}^3U_{i,n+1} - E & -\sqrt{3} V_{i,n+1} \\ V_{i,n+1} & -\sqrt{3} V_{i,n+1} & U_{n+1,i} - E \end{vmatrix} = 0 \quad (29)$$

where we have simplified the notation by dropping the subscript Σ for the $(R_{(i)}H_{(n+1)})^+$ fragment. Using the $(RH)^+$ potentials ${}^1W_{i,n+1}$ and ${}^3W_{i,n+1}$ (asymptotic to $R^+ + H$) and $W_{n+1,i}$ (asymptotic to $R + H^+$) as the eigenvalues $E^{(1)}$, $E^{(2)}$ and $E^{(3)}$ for Eq. (29), we can find numerically the diabatic potentials ${}^1U_{i,n+1}$, ${}^3U_{i,n+1}$ and $U_{n+1,i}$. The construction of the exchange term $V_{i,n+1}$ of the $(RH)^+$ diatomic fragment was described in Subsection II.A (see Eq. (4)). The diabatic Π -potentials ${}^1U_{i,n+j,\Pi}$ and ${}^3U_{i,n+j,\Pi}$ coincide with the adiabatic ${}^1W_{i,n+j,\Pi}$ and ${}^3W_{i,n+j,\Pi}$ potentials, since the exchange term $V_{i,n+j,\Pi}$ is equal to zero. The adiabatic diatomic potentials of the diatomic fragments R_2 , H_2 , RH , R_2^+ , H_2^+ and $(RH)^+$ for $R = Ar, Xe$ are presented in the Appendix.

III. Argon and Argon-Hydrogen Ionic Clusters Ar_n^+ , $(Ar_nH)^+$ and $(Ar_nH_2)^+$

Considering the structure of rare gas-hydrogen ionic clusters, we have decided to look at Ar and Xe, since they are expected to form different kinds of clusters with H due to the difference in their ionization potentials. Specifically, the Ar ionization potential ($I_{Ar} = 15.76$ eV) is much higher than the H ionization potential ($I_H = 13.6$ eV), whereas the Xe ionization potential ($I_{Xe} = 12.13$ eV) is smaller than that of H, which leads to different charge distributions in the argon-hydrogen and xenon-hydrogen clusters. But before considering the rare gas-hydrogen systems, we shall present the results of calculations on the rare gas ionic clusters Ar_n^+ and Xe_n^+ .

A. Ar_n^+ Clusters

The results of the calculations of the small ionic clusters Ar_3^+ , Ar_4^+ , Ar_5^+ and Ar_6^+ are presented in Table 1 and, partly, in Fig. 1. We shall now compare these results with the known experimental and theoretical results which have been obtained for the Ar_n^+ clusters.

According to experimental studies of the simplest cluster, Ar_3^+ , its energy of dissociation, $\text{Ar}_3^+ \rightarrow \text{Ar}_2^+ + \text{Ar}$, is about 0.2 eV.^{17,21,23,59,60} There are two alternative structures of the Ar_3^+ cluster, namely an asymmetrical Ar_2^+Ar structure with the charge concentrated on Ar_2^+ and a structure with the charge distributed among all three atoms. In the case of the Ar_2^+Ar structure, where the neutral atom is bound to the Ar_2^+ molecule by polarization forces, the triangle geometry is expected to be the most stable one.⁶¹ When the charge is delocalized among all three atoms, the exchange interaction contributes significantly to the binding so that the symmetrical linear geometry is expected to be the most stable configuration. The first quantum chemistry calculation, performed by the approximate X_α method, gave preference to the asymmetrical triangle structure.⁶² However, the CI ab initio calculations,^{24,25} as well as semiempirical calculations,^{27,29} found the symmetrical linear configuration to be the most stable one. The experimental spectral data, unfortunately, cannot provide direct evidence about the Ar_3^+ geometry. The conclusions made by the indirect analysis of the experimental data are contradictory. Whereas some of the experimental works confirm the symmetrical linear Ar_3^+ structure,^{17,35,63} other papers substantiate the asymmetrical Ar_2^+Ar structure, either of triangular⁶¹ or linear^{19,64} geometry.

Such discrepancy in the conclusions is due, most probably, from the features of the Ar_2^+ potential energy surface. Since Ar_3^+ has a very flat Ar_2^+-Ar potential,^{19,28,64} the vibrational motion shifts, for the most part, the Ar_3^+ configuration far away from the equilibrium configuration, thus significantly affecting the electronic transitions. This problem has been considered in detail by Gidea and Amarouche,²⁸ who calculated the potential energy surfaces of Ar_3^+ by a method similar to DIIS. A trajectory study performed on these potential energy surfaces shows that the symmetrical linear

Ar_3^+ minimum energy configuration alone provides the spectrum close to the experimental one, if one takes into account the vibrational motion.²⁸

In our calculations the triatomic cluster Ar_3^+ has a symmetrical linear equilibrium configuration, in accord with other calculations.²⁴⁻²⁸ The $\Sigma_k \rightarrow \Sigma_g$ transition energy in the equilibrium configuration is found in our calculations to be equal to 2.26 eV, close to the other theoretical values of 2.36 eV²⁴ and 2.34 eV.²⁸ Our $\Sigma_k \rightarrow \Sigma_g$ transition moment is 8.8 D, compared to the ab initio value of 8.24 D.²⁴ For the Ar-Ar distance at the equilibrium configuration, our calculation gives $R = 2.59 \text{ \AA}$, exactly the same as the semiempirical calculation of Gadea and Amarouche²⁸ but slightly less than ab initio calculations (2.62 \AA).²⁵ The $\text{Ar}_3^+ \rightarrow \text{Ar}_2^+ + \text{Ar}$ dissociation energy in our calculation is $D = 0.203 \text{ eV}$, close to the experimental values of 0.22 eV^{17,61} and $0.18 \pm 0.05 \text{ eV}$.²¹ Other calculations give values a little bit lower than our dissociation energies, namely, 0.18 eV,²⁴ 0.16 eV,²⁵ 0.20 eV²⁷ and 0.17 eV.²⁸

The most stable Ar_4^+ cluster, according to our calculations, is formed by an Ar_3^+ ion and almost neutral ($q = +0.002$) Ar atom separated from the nearest charged atom by 3.68 \AA (see Table 1 and Fig. 1.III). The dissociation energy for the $\text{Ar}_4^+ \rightarrow \text{Ar}_3^+ + \text{Ar}$ detachment process is found to be equal to $D = 0.047 \text{ eV}$, a typical energy of the polarization attraction. Another Ar_4^+ cluster is found in an asymmetrical linear configuration with one almost neutral ($q = +0.01$) atom separated from the nearest atom of Ar_3^+ by 3.3 \AA (see Table I and Fig. 1.V). The dissociation energy of this quasistable configuration is $D = 0.031 \text{ eV}$, and the potential barrier of the transition to the most stable configuration (Fig. 1.III) is about 0.001 eV only. In the DIM calculation,²⁷ the linear configuration with the dissociation energy of 0.043 eV is found to be the most stable one. According to experimental studies,^{20,65} the spectrum

of Ar_4^+ and other small Ar_n^+ clusters is similar to the Ar_3^+ spectrum, so that the conclusion is made that the Ar_n^+ ($n > 3$) clusters have a structure of $\text{Ar}_3^+\text{Ar}_{n-3}$. This experimental finding supports our result that Ar_3^+Ar is the most stable structure for the Ar_4^+ cluster. A similar result is obtained for Ar_5^+ and Ar_6^+ (see below). Like in the experimental studies, the spectrum of the most stable Ar_4^+ cluster is found to be similar to that of Ar_3^+ (Table 2.III). It consists of two transitions with large transition moments and transition energies close to one another and to the $^2\Sigma_g^+$ excitation energy of Ar_3^+ , so that they give one absorption band with its center close to that of the Ar_3^+ absorption. It is interesting to note that in the Ar_4^+ excitation states a strong charge transfer takes place, which makes it impossible to describe the excited $(\text{Ar}_4^+)^*$ cluster as an $(\text{Ar}_3^+)^*\text{Ar}$ structure (Table 2). Save for the transitions which resemble the $^2\Sigma_u^+ \rightarrow ^2\Sigma_g^+$ transition of Ar_3^+ , we find also in Ar_4^+ a low-energy (1.79 eV) transition with a small transition moment of $\mu = 0.7 \text{ D}$ (Table 2).

In addition to the Ar_4^+ clusters with the ground-state Ar_3^+Ar structure, we find a quasistable symmetrical configuration with all atoms bearing the same charge of $q = +0.25$, (Table 1.IV). In this configuration, Ar_4^+ has the geometry of a regular pyramid with a separation of 2.836 Å between every two atoms. The energy of this fully-symmetrical structure is only 0.1 eV above the energy of the most stable Ar_4^+ cluster. The spectrum of the symmetrical Ar_4^+ cluster is quite different from that of Ar_3^+ . The transition energy to the three lowest excited states which form a degenerate level is only 1.46 eV, compared to 2.26 eV for the first allowed transition in Ar_3^+ (Table 2). The symmetrical Ar_4^+ configuration is separated from the stable Ar_3^+Ar configuration by a high barrier of roughly 0.25 eV.

The Ar_5^+ and Ar_6^+ clusters are formed in a similar way as the stable Ar_4^+ cluster, i.e., by a Ar_3^+ ion and neutral or almost neutral Ar atoms. The energy of one Ar atom detachment is 0.05 eV for Ar_5^+ and 0.06 for Ar_6^+ (Table 1). The Ar_5^+ and Ar_6^+ spectra are similar to the spectrum of the stable Ar_4^+ cluster (Table 2.III). The experimental studies confirm that Ar_3^+ is an ionic core in small Ar_n^+ clusters.^{35,65} However, in large Ar_n^+ clusters, $n \geq 15$, the Ar_2^+ molecule was found to be the most stable ionic core.⁶⁵⁻⁶⁷ The transition from the Ar_3^+ to Ar_2^+ core with an increase of cluster size can be explained by the polarization effects. The energy of the neutral atoms polarization by the ionic is slightly larger for the Ar_2^+ than Ar_3^+ core because of the difference in their dimension. In small Ar_n^+ clusters, this difference in polarization energy is small compared to the Ar_3^+ dissociation energy, and consequently it is not expected to affect the core structure. In large Ar_n^+ clusters, the polarization energy is close to that in solid Ar where, according to our estimations, the difference between the polarization energies of the Ar_2^+ and Ar_3^+ cores lies somewhere between 0.3 and 0.6 eV. This energy difference is larger than the Ar_3^+ dissociation energy, so that the Ar_2^+ core may become more stable than the Ar_3^+ one.

B. $(\text{Ar}_n\text{H})^+$ Clusters

The structure of the $(\text{ArH})^+$ clusters is simple: all of them are formed by an $(\text{ArH})^+$ ionic molecule and $n-1$ neutral Ar atoms which are bound to $(\text{ArH})^+$ by polarization forces (Table 3). Since in the $(\text{ArH})^+$ molecule the H atom bears more charge than the Ar atom ($q_{\text{H}} = +0.569$), the neutral Ar atoms are located on the H side of the $(\text{ArH})^+$ ion (Fig. 2). The simplest $(\text{Ar}_n\text{H})^+$ cluster, $\text{Ar}(\text{ArH})^+$, has a bent geometry with the neutral Ar atom at a distance of 2.81 Å from H. In the $\text{Ar}_2(\text{ArH})^+$, $\text{Ar}_3(\text{ArH})^+$ and $\text{Ar}_4(\text{ArH})^+$ clusters, the

neutral atoms are located symmetrically around the $(\text{ArH})^+$ molecular axis at the distances of 2.79 Å, 2.79 Å and 2.84 Å, respectively. In the $\text{Ar}_6(\text{ArH})^+$ clusters, two of the Ar atoms are more tightly bound to $(\text{ArH})^+$ than the other four Ar atoms. In the sequence of $\text{Ar}_k(\text{ArH})^+$ clusters, the H charge increases monotonically with k . The energy of an Ar atom detachment from an $\text{Ar}_k(\text{ArH})^+$ cluster is relatively large, varying from 0.09 eV for $k = 6$ to 0.19 eV for $k = 3$. The electronic spectrum of the $\text{Ar}_k(\text{ArH})^+$ clusters is practically the same as the spectrum of the isolated $(\text{ArH})^+$ molecule. The $\text{Ar}_k(\text{ArH})^+$ clusters can probably be detected by studying their rotational spectrum, since the $\text{Ar}_k(\text{ArH})^+$ dipole moment depends strongly on k (Table 3).

C. $(\text{Ar}_n\text{H}_2)^+$ Clusters

The ionization potential of Ar (15.76 eV) is close to the ionization potential of the H_2 molecule (15.426 eV), which leads to strong coupling between the states Ar^+H_2 and ArH_2^+ . This coupling affects significantly the potential energy surfaces of the $(\text{ArH}_2)^+$ system, which have been calculated by the DIM method.⁴⁷ We shall not consider here the $(\text{ArH}_2)^+$ potential energy surfaces, i.e., we shall not consider the energy E as a function of the interatomic distances $R_{\text{Ar-H}(1)}$, $R_{\text{Ar-H}(2)}$ and $R_{\text{H-H}}$, but restrict our task to the study of ionic clusters formed by the Ar atom or atoms and the H_2 molecule. The calculation of the $(\text{Ar}_n\text{H}_2)^+$ system is performed by the modified DIIS method described in Subsection II.C above.

The $(\text{ArH}_2)^+$ clusters can be formed by $\text{Ar}^+ + \text{H}_2$ (-4.746 eV), $\text{Ar} + \text{H}_2^+$ (-4.95 eV) and $\text{H} + (\text{ArH})^+$ (-6.215 eV). It follows that the most stable $(\text{ArH}_2)^+$ cluster is expected to be formed by the $(\text{ArH})^+$ molecule and H atom. However, the polarization attraction of the H atom to the $(\text{ArH})^+$ ion is so

weak that, according to our calculation, no $(\text{ArH}_2)^+$ cluster can be formed. The ionic $(\text{ArH}_2)^+$ cluster can be formed in a quasistable configurations only. We find that the quasistable $(\text{ArH}_2)^+$ cluster with the lowest energy is formed by an almost neutral Ar atom ($q = 0.02$) and the H_2^+ molecular ion with an H-H separation of 1.05 Å, slightly shorter than in the isolated H_2^+ (1.06 Å). The Ar atom is located in a plane perpendicular to the H_2^+ axis at a distance of 2.935 Å from the H_2^+ center (Fig. 2). The energy of the Ar detachment is relatively large, $D = 0.193$ eV. The ArH_2^+ cluster is separated from the stable state of separated $(\text{ArH})^+$ and H by a high barrier of, roughly, 0.75 eV. The saddle point for the $\text{ArH}_2^+ \rightarrow \text{H} + (\text{ArH})^+$ dissociation lies in the collinear configuration with the interatomic distances $R_{\text{H-H}} \approx 1.5$ Å and $R_{\text{Ar-H}} \approx 1.2$ Å. Unfortunately, in the DIM calculation⁴⁷ the geometry of Fig. 2 was not studied so we do not have a comparison for our results.

The Ar_nH_2^+ clusters with a few Ar atoms have a structure similar to that of ArH_2^+ (see Fig. 2 and Table 4). In all of these clusters up to $n = 5$, the Ar atoms are located in a plane perpendicular to the H_2^+ axis at distances of about 3 Å. The energy of the Ar detachment is 0.19-0.2 eV, slowly decreasing with n for $n > 2$. The common charge of the Ar atoms slightly increases with n reaching $q = \pm 0.075$ in Ar_5H_2^+ . The decrease of the H_2^+ charge with n leads to some decrease of the H-H interatomic distance, which is 1.03 Å in Ar_5H_2^+ . The larger Ar_nH_2^+ clusters ($n > 5$) are expected to be much less stable than the clusters presented in Table 4. The photoabsorption spectrum of the Ar_nH_2^+ clusters is very different from the H_2^+ spectrum, since the lowest excited states are built up by charge transfer to Ar atoms and the formation of Ar_n^+H_2 (Table 4).

IV. Xenon and Xenon-Hydrogen Clusters Xe_n^+ , $(\text{Xe}_n\text{H})^+$ and $(\text{Xe}_n\text{H}_2)^+$

A. Xe_n^+ Clusters

The results of the calculations of the small ionic clusters Xe_3^+ , Xe_4^+ , Xe_5^+ and Xe_6^+ are presented in Table 5 and (partly) in Fig. 3. In our calculations, the triatomic cluster Xe_3^+ has a symmetrical linear structure with 50% of the charge at the central atom, in accord with other calculations.^{24,29} Our calculation gives 3.38 Å for the Xe-Xe separation in Xe_3^+ , compared to 3.47 Å²⁴ and 3.32 Å²⁹ in other calculations. The $\text{Xe}_3^+ \rightarrow \text{Xe}_2^+ + \text{Xe}$ dissociation energy in our calculation is $D = 0.20$ eV, which is 0.07 eV lower than the experimental value of 0.27 eV.²² Our dissociation energy is almost exactly halfway between the values of 0.12 eV²⁴ and 0.36 eV²⁶ obtained in two other calculations. For the $^2\Sigma_u^+ \rightarrow ^2\Sigma_g^+$ transition energy from the equilibrium point, we obtain $T = 1.87$ eV, a little bit higher value than in Ref. 24 (1.60 eV).

Our results obtained for Xe_n^+ clusters with $n > 3$ are quite different from those obtained in Ref. 29. According to our calculations, the most stable Xe_4^+ cluster has a structure of a regular pyramid, like Ar_4^+ (Fig. 1.IV), and the lowest energy quasistable Xe_4^+ cluster has a structure Xe_3^+Xe with the weakly-charged atom ($q = 0.014$) located to the side of the Xe_3^+ axis (Table 5, Fig. 3). We find also a quasistable linear symmetrical Xe_4^+ cluster which energy lies 0.012 eV above the Xe_3^+Xe . In contrast to this, in Ref. 29 the most stable Xe_4^+ structure (Xe detachment energy of 0.159 eV) is just the linear symmetrical one. A structure like our regular pyramid (Fig. 3.III) was not found in Ref. 29 at all. We find the collinear Xe_4^+ cluster V to be separated from the bent geometry Xe_3^+Xe cluster IV by a small barrier of 0.03 eV. The Xe_3^+Xe cluster IV which can be formed easily as the result of the attachment of Xe to Xe_3^+ (attachment energy of 0.08 eV) is separated by a high

barrier of approximately 0.3 eV from the most stable Xe_4^+ cluster, III. It follows that the Xe_3^+X cluster can exist for a relatively long time when the temperature is not high.

The clusters Xe_5^+ and Xe_6^+ are presented in Table 5 for both the $\text{Xe}_4^+\text{Xe}_{n-4}$ and $\text{Xe}_3^+\text{Xe}_{n-3}$ structures. The energy of Xe detachment varies in the interval 0.07-0.11 eV. The binding energy of the attached atoms is derived from polarization forces, since these atoms are almost neutral.

In the family of $\text{Xe}_3^+\text{Xe}_{n-3}$ clusters (IV, VII and IX in Table 5), the strongest transition is to the excited state similar to $^2\Sigma_g^+$ of Xe_3^+ , i.e., with a very small charge on the central Xe_3^+ atom. In some of these clusters, an important portion of the charge is transferred in this excited state from Xe_3^+ to the attached atoms, for example, $q = 0.38$ in the Xe_3^+ cluster IV. As in the case of the $\text{Ar}_3^+\text{Ar}_{n-3}$ clusters, the transition energy is not affected much by the charge transfer, so that in the $\text{Xe}_3^+\text{Xe}_{n-3}$ clusters a strong photoabsorption is expected in the same range as in Xe_3^+ . In addition to this $\Sigma_u \rightarrow \Sigma_g$ -like transition, more allowed transitions with smaller transition energy are found in the $\text{Xe}_3^+\text{Xe}_{n-3}$ clusters (Table 6).

In the symmetrical Xe_4^+ configuration III, there is threefold degenerate level with an energy of 1.20 eV above the ground state (and a transition moment of 2.7 D). In Xe_4^+Xe and Xe_4^+Xe_2 , two of the excited states remain degenerate with almost the same excitation energy as in Xe_4^+ , whereas the third state excitation energy is lowered by about 0.15 eV. The linear Xe_4^+ cluster has only one strong transition $\Sigma_u \rightarrow \Sigma_g$, whose energy is 1.46 eV, about 0.4 eV less than in Xe_3^+Xe . The differences between the spectra of Xe_4^+ III (regular pyramid), Xe_3^+Xe IV and Xe_4^+ V (linear) may help in the experimental search for the most stable Xe_4^+ cluster.

B. $(\text{Xe}_n\text{H})^+$ Clusters

As in the case of the $(\text{Ar}_n\text{H})^+$ clusters, the $(\text{Xe}_n\text{H})^+$ clusters are formed by the $(\text{XeH})^+$ molecule which attracts more neutral Xe atoms. Since in the $(\text{XeH})^+$ molecule the Xe atom bears most of the charge, the neutral Xe atoms are located on the rare gas atom side of $(\text{XeH})^+$ (Fig. 4). The $\text{Xe}(\text{XeH})^+$ cluster is found to have a linear geometry with a $\text{Xe} + (\text{XeH})$ dissociation energy of 0.046 eV, which is much less than the $\text{Ar}(\text{ArH})^+$ dissociation energy (0.163 eV). The energy of the Xe detachment in the $\text{Xe}_k(\text{XeH})^+$ clusters varies in the interval 0.04-0.08 eV (Table 7). The charge of H in $(\text{XeH})^+$ slightly decreases when the number of attached neutral atoms increases. The spectrum of the $\text{Xe}_k(\text{XeH})^+$ clusters is practically the same as that of isolated $(\text{XeH})^+$.

C. $(\text{Xe}_n\text{H}_2)^+$ Clusters

The most stable $(\text{Xe}_n\text{H}_2)^+$ clusters are expected to be formed by charged Xe atoms and a neutral H_2 molecule, since the H_2 ionization potential is about 3.3 eV higher than that of Xe. The calculation of the simplest, $(\text{XeH}_2)^+$, system shows the existence of a strongly-bound cluster with the energy of the $\text{Xe}^+ + \text{H}_2$ dissociation of $D = 0.456$ eV. This cluster has a linear geometry with the Xe-H distance of 1.79 Å, 0.18 Å larger than in the $(\text{XeH})^+$ molecule, and the H-H distance of 0.81 Å, 0.07 Å larger than in H_2 molecule (Fig. 4). Almost a whole charge is concentrated on the Xe atom ($q = 0.955$), but the valence bond between Xe^+ and H_2 is strong. The excitation energies for the allowed transitions (in the equilibrium configuration) are 10.76 eV and 18.56 eV, whereas in H_2 there is one transition, 19.2 eV, and in $(\text{XeH})^+$ there are two transitions but with energies of 7.93 eV and 13.3 eV.

The structure of $(\text{Xe}_n\text{H}_2)^+$ clusters with more than one Xe atom is quite different from that of $(\text{XeH}_2)^+$ due to the formation of Xe_2^+ or Xe_3^+ molecules.

All $(\text{Xe}_n\text{H}_2)^+$ clusters that we have studied ($n > 1$) have the structure Xe_n^+H_2 with a practically neutral H_2 molecule. In the most stable Xe_2^+H_2 cluster, the axis of both molecules are perpendicular to one another (Fig. 4). The dissociation energy of this cluster is 0.084 eV. In the Xe_3^+H_2 clusters three Xe atoms form, in contrast to the free Xe_3^+ cluster, a triangular Xe_3^+ system with Xe-Xe distances in the interval 3.50-3.58 Å. In one of the Xe_3^+H_2 isomers with the $\text{Xe}_3^+\text{H}_2 \rightarrow \text{Xe}_3^+ + \text{H}_2$ dissociation energy of $D = 0.108$ eV, the Xe_3^+ plane is perpendicular to the H_2 axis with its center located on this axis at 3.08 Å from the H_2 center. In another isomer ($D = 0.097$ eV), H_2 is located in the Xe_3^+ plane. In contrast to the $\text{Xe}_k(\text{XeH})^+$ clusters, in Xe_n^+H_2 clusters ($n \geq 2$), the valence binding affects to a certain degree the cluster structure.

V. Conclusions

1. The semiempirical diatomics-in-molecules (DIIS) method⁹ proves to be an effective tool to treat rare gas R_n^+ and rare gas-hydrogen $(\text{R}_n\text{H})^+$ ionic clusters. The DIIS results obtained for the triatomic Ar_3^+ and Xe_3^+ clusters are close to those obtained in other studies.

2. According to our calculation, there are two very different isomers of R_4^+ clusters. One has the structure of a R_3^+ ion and almost neutral R atom, whereas the other isomer has the symmetrical geometry of a regular pyramid with all four atoms bearing the same charge. We find the Ar_3^+Ar and the symmetrical Xe_4^+ clusters to be the most stable isomers, so that Ar_n^+ ($n = 4-6$) and Xe_n^+ ($n = 5,6$) to have ionic cores Ar_3^+ and Xe_4^+ , respectively. The experimentally detected transition from the Ar_3^+ to Ar_2^+ ionic core in large ($n \geq 15$) Ar_n^+ clusters results, most probably from the polarization energy increase.

3. The $(R_n H)^+$ clusters have the structure of a $(RH)^+$ molecule and $n-1$ neutral R atoms. The neutral rare gas atoms are located at the H side of the $(ArH)^+$ molecule in the $(ArH)^+ Ar_{n-1}$ clusters and at the Xe side of the $(XeH)^+$ molecule in the $(XeH)^+ Xe_{n-1}$ clusters.

4. In order to consider the $(R_n H_2)^+$ clusters, we have modified the DIIS method by including into the DIIS wave function some DIM components which describe $^3\Sigma$ and $^3\Pi$ states of the R^+H fragments. We find that the Xe^+ ion forms with H_2 molecule a stable ($D_e = 0.46$ eV) collinear cluster $(XeHH)^+$ with the strong contribution of valence forces. When more Xe atoms are involved, the cluster is formed by Xe_n^+ ($n > 1$) ion and H_2 molecule which are bound one to another by polarization forces only. In the case of Ar atoms, we do not find any stable ionic $(ArH_2)^+$ clusters. The quasistable clusters are formed by slightly-charged Ar atoms and H_2^+ ionic molecule.

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Appendix: Diatomic Potentials and Exchange Terms

The matrix elements within the DIIS method are expressed in terms of the potentials and exchange terms of diatomic fragments (see Sec. II). In the R_n^+ , $(R_n H)^+$ and $(R_n H_2)^+$ ionic clusters considered in this work, there are the following neutral and ionic diatomic fragments: R_2 , H_2 , RH , R_2^+ , H_2^+ and $(RH)^+$. The potentials of the neutral fragments, W_{ij} , are used directly in the matrix elements expressions. The adiabatic potentials of the ionic fragments,

W_{ij} , are needed for the calculations of the diabatic potentials U_{ij} (see Eqs. (3), (4) and (29)). When the heteronuclear fragment $(RH)^+$ is considered, one needs also the exchange term V which has to be defined independently from the potentials. The potentials are given in eV and the distances in Å.

A. Neutral Fragments R_2 , H_2 and RH

We need the following potentials of the neutral fragments: $^1\Sigma_{RR}$, $^1\Sigma_{HH}$, $^1\Sigma_{RH}$ and $^3\Sigma_{HH}$ (the last one is for $(R_nH_2)^+$ only). The R_2 potentials are taken as a sum of a repulsive exponential term and an attractive dispersion term,

$$W_{ArAr}(^1\Sigma) = 6940 \exp[-3.6R] - 60/R^6, \quad (A.1)$$

which provides the empirical⁶⁸ values $R_e = 3.76$ Å and $D_e = 0.012$ eV and fits the ab initio potential⁶⁹ at small R , and

$$W_{XeXe}(^1\Sigma) = 24000 \exp[-3.17R] - 320/R^6, \quad (A.2)$$

which provides the empirical⁷⁰ values $R_e = 4.47$ Å and $D_e = 0.023$ eV and fits the ab initio potential⁷¹ at small R .

The ground-state H_2 potential is described by a modified (for large R) Morse function,⁵⁸

$$W_{HH}(^1\Sigma_g) = 4.746 G(G-2)g \quad (A.3)$$

$$G = \exp[-a(R-R_e)] \quad (A.4)$$

$$g = \exp[-b(R-R_e)^3] \quad (A.5)$$

In Eqs. (A.4) and (A.5), $R_e = 0.7417 \text{ \AA}$, and the parameters $a = 1.9446 \text{ \AA}^{-1}$ and $b = 0.1215 \text{ \AA}^{-3}$ are found by fitting the exact H_2 potential.⁷² The accuracy of the fit is generally within 0.05 eV for $R > 0.53 \text{ \AA}$. The repulsive $^3\Sigma$ potential of H_2 is described by an anti-Morse function,

$$W_{HH}(^3\Sigma) = 1.963 G(G-2)g, \quad (A.6)$$

where the G-function parameter is $a = 1.805 \text{ \AA}^{-1}$ and the g-function parameter is $b = 0.1215 \text{ \AA}^{-3}$. The potential (A.6) fits the exact potential.⁷²

The RH potential is described by a Lennard-Jones potential with the empirical parameters for ArH of $R_e = 3.62 \text{ \AA}$ and $D_e = 4.15 \text{ meV}$ ⁷³ and for XeH of $R_e = 3.93 \text{ \AA}$ and $D_e = 6.85 \text{ meV}$:⁷⁴

$$W_{ArH}(^1\Sigma) = 0.0166 \left[\left(\frac{3.22}{R} \right)^{12} - \left(\frac{3.22}{R} \right)^6 \right] \quad (A.7)$$

$$W_{XeH}(^1\Sigma) = 0.0274 \left[\left(\frac{3.5}{R} \right)^{12} - \left(\frac{3.5}{R} \right)^6 \right] \quad (A.8)$$

B. Ionic Homonuclear Fragments R_2^+ and H_2^+

We need the following adiabatic potentials of the ionic homonuclear fragments: $^2\Sigma_u$, $^2\Sigma_g$, $^2\Pi_g$, $^2\Pi_u$ for R_2^+ and σ_g , σ_u for H_2^+ . The attractive $^2\Sigma_u$ and $^2\Pi_g$ potentials of R_2^+ are described as a sum of Morse function and the polarization term P . For argon these are expressed as

$$W_{(ArAr)}(^2\Sigma_u) = 1.184 G(G-2) + P_{Ar}, \quad R_e = 2.48, \quad a = 2.0 \quad (A.9)$$

$$W_{(ArAr)}(^2\Pi_g) = 0.025 G(G-2) + P_{Ar}, \quad R_e = 3.2, \quad a = 2.3 \quad (A.10)$$

$$P_{\text{Ar}} = -11.7 \left[1 + \left(\frac{3.8}{R} \right)^{12} \right]^{-1/3} / R^4, \quad (\text{A.11})$$

where R_e and a in (A.9) and (A.10) are the parameters of the G-function (A.4), and the number 3.8 in (A.11) is twice the radius of the Ar atom. We assume the Ar^+ radius to be equal to the neutral Ar radius $r_{\text{Ar}} = 1.9 \text{ \AA}$. For xenon we have

$$W_{(\text{XeXe})}^{+(^2\Sigma_u)} = 1.0 G(G-2) + P_{\text{Xe}}, \quad R_e = 3.22, \quad a = 155 \quad (\text{A.12})$$

$$W_{(\text{XeXe})}^{+(^2\Pi_g)} = 0.01 G(G-2) + P_{\text{Xe}}, \quad R_e = 3.8, \quad a = 3.0 \quad (\text{A.13})$$

$$P_{\text{Xe}} = -2.91 \left[1 + \left(\frac{4.4}{R} \right)^{12} \right]^{-1/3} / R^4. \quad (\text{A.14})$$

The Xe and Xe^+ radii are both taken as equal to $r_{\text{Xe}} = 2.2 \text{ \AA}$, so that twice the double radius in (A.14) is 4.4 \AA .

The repulsive potentials of R_2^+ are described by a sum of a polarization term P and anti-Morse function for the Σ_g state or exponential function for the Π_u state:

$$W_{(\text{ArAr})}^{+(^2\Sigma_g)} = 0.9 G(G+2) + P_{\text{Ar}}, \quad R_e = 2.48, \quad a = 1.72 \quad (\text{A.15})$$

$$W_{(\text{ArAr})}^{+(^2\Pi_u)} = 1290 \exp[-2.8R] + P_{\text{Ar}} \quad (\text{A.16})$$

$$W_{(\text{XeXe})}^{+(^2\Sigma_g)} = 0.78 G(G+2) + P_{\text{Xe}}, \quad R_e = 3.22, \quad a = 1.45 \quad (\text{A.17})$$

$$W_{(\text{XeXe})}^{+(^2\Pi_u)} = 6.910 \exp[-2.75R] + P_{\text{Xe}}. \quad (\text{A.18})$$

The ground-state potentials (A.9) and (A.12) provide the empirical dissociation energies⁷⁵ of $D_e = 1.24$ eV for Ar_2^+ and $D_e = 1.08$ eV for Xe_2^+ . The excited-state potentials (A.10), (A.13) and (A.15)-(A.18) fit the ab initio potentials of Ref. 75.

The exact ground-state potential σ_g of the H_2^+ molecular ion⁷⁶ has been fit by a Morse function in Ref. 77 (there is a misprint in Ref. 77 on p. 17: $a = 0.6678$ should be $a = 0.678$). In order to improve the fitting at small R , we have added a repulsive exponential component to the Morse function. At large R , the fitting is improved by the g -function (A.5),

$$W_{(\text{HH})^+}(\sigma_g) = 810.7 \exp[-11.34R] + 2.795 G(G-2)g \quad (\text{A.19})$$

with the following parameters for the G - and g -functions: $R_e = 1.054$ Å, $a = 1.281$ Å, $b = 0.0067$ Å⁻³. The accuracy of the fit is within 0.015 eV for $R > 0.74$ Å. The equilibrium distance of the potential (A.19) is 1.06 Å, slightly larger than the parameter R_e for the G - and g -functions. The dissociation energy of the potential (A.19) is $D_e = 2.79$ eV. The excited σ_u potential⁷⁶ is fit by an anti-Morse potential,

$$W_{(\text{HH})^+}(\sigma_u) = 3.014 G(G+2)g, \quad R_e = 1.054, \quad a = 1.379, \quad b = 0.0067. \quad (\text{A.20})$$

The polarization component of the potentials (A.19) and (A.20) is

$$P_H = -4.8 \left[1 + \left(\frac{1.7}{R} \right)^{12} \right]^{-1/3} / R^4, \quad (\text{A.21})$$

where 1.7 is the H radius.

C. Ionic Heteronuclear Fragment (RH)⁺

In order to find from Eq. (4) or (29) the (RH)⁺ diabatic Σ potentials U, one needs three adiabatic potentials, namely $^1\Sigma$ and $^3\Sigma$ asymptotic to R⁺ + H and $^1\Sigma$ asymptotic to R + H⁺. The adiabatic potentials are described as a sum of an attractive (Morse) or repulsive (anti-Morse or exponential) potential and the polarization term P. The adiabatic Σ potentials of (ArH)⁺ are expressed as follows:

$$W_{(\text{ArH})+(X^1\Sigma)} = 3.157 G(G-2) + P_{\text{Ar}} \quad , \quad R_e = 1.266 \quad , \quad a = 1.94 \quad (\text{A.22})$$

$$W_{(\text{ArH})+(B^1\Sigma)} = 2.75 G(G+2) + P_{\text{H}} + 2.16 \quad , \quad R_e = 1.266 \quad , \quad a = 1.2 \quad (\text{A.23})$$

$$W_{(\text{ArH})+(^3\Sigma)} = 600 \exp[-1.8R] + P_{\text{H}} + 2.16 \quad (\text{A.24})$$

$$P_{\text{Ar}} = 11.7[1 + (\frac{1.9}{R})^{12}]^{-1/3}/R^4 \quad (\text{A.25})$$

$$P_{\text{H}} = 4.8[1 + (\frac{3.6}{R})^{12}]^{-1/3}/R^4 \quad (\text{A.26})$$

In (A.22) and (A.23), R_e and a are the parameters for the G-function (A.4), and in (A.23) and (A.24) the number 2.16 presents the difference between the ionization potentials of Ar and H. In the polarization terms (A.25) and (A.26), 1.9 is the Ar radius and 3.6 is the sum of the Ar and H radii. The ground-state (ArH)⁺ equilibrium distance $R_e = 1.266 \text{ \AA}$ in the potential (A.22) is equal to the experimental value.^{78,79} This potential provides the ab

initio dissociation energy $D_e = 4.055$ eV.⁸⁰ The ab initio ground-state potential of Ref. 80 is used to find the parameter a of (A.22). The excited-state potentials (A.23) and (A.24) are found by fitting the ab initio potentials of Ref. 81. The $(\text{ArH})^+$ exchange term V is found from Eq. (4) by fitting the ab initio dipole moment,⁸⁰ which is close to the experimental value⁸¹ at the equilibrium point,

$$V = -95.5 R^2 \exp[-2.42R] \quad (\text{A.27})$$

The adiabatic Σ potentials of $(\text{XeH})^+$ are as follows:

$$W_{(\text{XeH})^+}(\text{X}^1\Sigma) = 4.03 G(G-2) + P_H, \quad R_e = 1.61, \quad a = 1.56 \quad (\text{A.28})$$

$$W_{(\text{XeH})^+}(\text{B}^1\Sigma) = 73 \exp[-1.3R] + P_{\text{Xe}} + 1.47 \quad (\text{A.29})$$

$$W_{(\text{XeH})^+}(\text{X}^3\Sigma) = 1.3 G(G+2) + P_H, \quad R_e = 1.61, \quad a = 1.56 \quad (\text{A.30})$$

$$P_H = -4.8[1 + (\frac{3.9}{R})^{12}]^{-1/3}/R^4 \quad (\text{A.31})$$

$$P_{\text{Xe}} = -2.91[1 + (\frac{2.2}{R})^{12}]^{-1/3}/R^4 \quad (\text{A.32})$$

In (A.29), 1.47 presents the difference between the ionization potentials of H and Xe. In the polarization terms (A.31) and (A.32), 3.9 is the sum of the Xe and H radii and 2.2 is the Xe radius. The ground-state $(\text{XeH})^+$ equilibrium distance $R_e = 1.61$ Å in the potential (A.28) is equal to the experimental value.⁸² The potential (A.28) provides the ab initio⁸³ dissociation energy $D_e = 4.05$ eV, a value close to the experimental dissociation energy of $D_e \approx 3.8$ eV.⁸⁴ The parameter a of (A.28) is found by fitting the ab initio potential of Ref. 83. The excited-state potentials (A.29) and (A.30) are found by

fitting the ab initio potentials of Ref. 85. The $(\text{XeH})^+$ exchange term V is found from Eq. (4) by fitting the ab initio dipole moment,⁸³

$$V = -57.7 R^2 \exp[-2.0 R] \quad . \quad (\text{A.33})$$

The Π potentials of $(\text{ArH})^+$ are obtained by fitting the ab initio potentials of Ref. 81,

$$W_{(\text{ArH})^+}(\text{A}^1\Pi) = 423 \exp[-3.9R] + P_H + 2.16 \quad (\text{A.34})$$

$$W_{(\text{ArH})^+}(\text{A}^3\Pi) = 90 \exp[-1.2R] + P_H + 2.16 \quad , \quad (\text{A.35})$$

where P_H is given by (A.26). The Π potentials of $(\text{XeH})^+$ are obtained by fitting the ab initio potentials of Ref. 85,

$$W_{(\text{XeH})^+}(\text{A}^1\Pi) = 370 \exp[-2.74R] + P_H \quad (\text{A.36})$$

$$W_{(\text{XeH})^+}(\text{A}^3\Pi) = 430 \exp[-2.74R] + P_H \quad , \quad (\text{A.37})$$

where P_H is given by (A.31).

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Table 1. Structure of the ionic clusters Ar_n^+ and ionic molecule Ar_2^+ . The distances (R) are in Å, and the energies (E,D) in eV.

N	Ar_n^+	Structure ¹⁾	R ²⁾	R ³⁾	q ⁴⁾	E	Dissociation Products	D
I	Ar_2^+		2.48			-1.240	$\text{Ar}^+ + \text{Ar}$	1.240
II	Ar_3^+	linear	2.59			-1.443	$\text{Ar}_2^+ + \text{Ar}$	0.203
III	Ar_4^+	$(\text{Ar}_3^+)_z(\text{Ar})_{xz}$	2.59	3.68	0.002	-1.490	$\text{Ar}_3^+ + \text{Ar}$	0.047
IV	Ar_4^+	regular pyramid	2.836			-1.480	$\text{Ar}_2^+ + 2\text{Ar}$	0.240
V	Ar_4^+	linear	2.59	3.3	0.01	-1.474	$\text{Ar}_3^+ + \text{Ar}$	0.031
VI	Ar_5^+	$(\text{Ar}_3^+)_z(\text{Ar}_2)_{xy}$	2.59	3.68	0	-1.540	$(\text{Ar}_4^+)_{\text{III}} + \text{Ar}$	0.050
VII	Ar_5^+	$(\text{Ar}_3^+)_z(\text{Ar}_2)_{xz}$	2.59	3.61	0.002	-1.540	$(\text{Ar}_4^+)_{\text{III}} + \text{Ar}$	0.050
VIII	Ar_5^+	$(\text{Ar}_3^+)_z(\text{Ar}_2)_{xz}$	2.59	3.68	0.003	-1.537	$(\text{Ar}_4^+)_{\text{III}} + \text{Ar}$	0.047
IX	Ar_6^+	$(\text{Ar}_3^+)_z(\text{Ar}_2)_{xy}(\text{Ar})_x$	2.59	3.61	0.001	-1.601	$(\text{Ar}_5^+)_{\text{VI}} + \text{Ar}$	0.061

¹⁾ The subscripts z, xz, etc. denote the location along an axis or in a plane (see Fig. 1.II-V).

²⁾ Distance between adjacent charged atoms.

³⁾ Distance from a neutral or weakly-charged atom to the nearest charged atom.

⁴⁾ Common charge of all weakly-charged atoms.

Table 2. Transition energy T in eV and transition moment μ in Debyes of the allowed transitions in Ar_3^+ and Ar_4^+ . q is the charge distribution in the excited states.

N ¹⁾	Ar_n^+	T	D	q			
				1	2	3	4
II	Ar_3^+	2.26	8.8	0.5	0	0.5	
III	Ar_3^+Ar	2.24	6.8	0.30	0.01	0.34	0.35
		2.19	2.3	0.10	0.04	0.24	0.62
		1.79	0.7	0.27	0.37	0.15	0.21
V	Ar_4^+ linear	1.97	8.2	0.24	0.01	0.26	0.49
IV	Ar_4^+ pyramid	1.46	2.1	0.25	0.25	0.25	0.25
		1.46	2.1	0.30	0.30	0.20	0.20
		1.46	2.1	0.20	0.20	0.30	0.30

¹⁾ For N, see Table 1.

Table 3. Structure of the ionic clusters $\text{Ar}_k(\text{ArH})^+$ and ionic molecule $(\text{ArH})^+$.

The distances (R) are in Å, the energies (E) in eV and the dipole moment (μ) in Debyes.

$\text{Ar}_k(\text{ArH})^+$	$q_{\text{H}}^{1)}$	$R_{\text{Ar-H}}^{2)}$	$R_{\text{Ar-Ar}}^{2)}$			μ	E	Ar Detachment
$(\text{ArH})^+$	0.569					2.78	-4.055	
$\text{Ar}(\text{ArH})^+$	0.579	2.81		3.70		5.98	-4.218	0.163
$\text{Ar}_2(\text{ArH})^+$	0.591	2.79		3.81		6.32	-4.395	0.177
$\text{Ar}_3(\text{ArH})^+$	0.601	2.79		3.72		6.23	-4.586	0.191
$\text{Ar}_4(\text{ArH})^+$	0.609	2.84		3.57		4.94	-4.760	0.174
$\text{Ar}_6(\text{ArH})^+$	0.614	2.84	3.17	3.87	3.38	3.33	-4.947	0.094

1) Charge of the H atom in $(\text{ArH})^+$

2) $R_{\text{Ar-H}}$ and $R_{\text{Ar-Ar}}$ are the distances from the neutral atom to H and Ar of $(\text{ArH})^+$, respectively. For $\text{Ar}_6(\text{ArH})^+$ there are two different Ar-H and Ar-Ar distances.

Table 4. Structure of the ionic clusters Ar_nH_2^+ and ionic molecule H_2^+ . The distances (R) are in Å, the energies in eV and the dipole moments (μ) in Debyes.

Ar_nH_2^+	$q_{\text{H}_2^+}$	$R_{\text{H-H}}$	R_{Ar}	μ	E	Ar Detachment	T	Transitions		
								μ	T	μ
H_2^+	1.0	1.06		0	-2.790		11.8	2.5		
ArH_2^+	0.98	1.05	2.935	12.3	-2.983	0.193	5.65	1.9		
Ar_2H_2^+	0.964	1.05	2.94	9.02	-3.183	0.200	5.83	2.0		
Ar_3H_2^+	0.95	1.04	2.95	5.15	-3.377	0.194	5.58	0.8	5.89	2.3
Ar_4H_2^+	0.936	1.03	2.955	1.65	-3.568	0.191	5.71	0.9	5.94	2.5
Ar_5H_2^+	0.926	1.03	2.995	0	-3.754	0.186	5.44	1.7	5.92	2.2

¹⁾ Distance from the Ar atoms to the center of H_2^+ . The adjacent Ar atoms are separated by 3.7 Å ($n = 2-4$) and 3.52 Å in symmetrical Ar_5H_2^+ .

Table 5. Structure of the ionic clusters Xe_n^+ and ionic molecule Xe_2^+ . (See Table 1 for the footnotes.)

N	Xe_n^+	structure ¹⁾	R ²⁾	R ³⁾	q ⁴⁾	E	Dissociation Products	D
I	Xe_2^+		3.22			-1.080	$\text{Xe}^+ + \text{Xe}$	1.08
II	Xe_3^+	linear	3.383			-1.277	$\text{Xe}_2^+ + \text{Xe}$	0.197
III	Xe_4^+	regular pyramid	3.646			-1.480	$\text{Xe}_2^+ + 2\text{Xe}$	0.40
IV	Xe_4^+	$(\text{Xe}_3^+)_z (\text{Xe})_{xz}$	3.384	4.19	0.014	-1.356	$\text{Xe}_3^+ + \text{Xe}$	0.079
V	Xe_4^+	linear	3.315	3.65	0.184	-1.346	$\text{Xe}_3^+ + \text{Xe}$	0.069
VI	Xe_5^+	$(\text{Xe}_4^+)_{\text{III}} (\text{Xe})_{xz}$	3.646	4.40	0.001	-1.568	$\text{Xe}_4^+ + \text{Xe}$	0.088
VII	Xe_5^+	$(\text{Xe}_3^+)_z (\text{Xe}_2)_{xz}$	3.383	4.24	0.005	-1.439	$\text{Xe}_3^+ \text{Xe} + \text{Xe}$	0.083
VIII	Xe_6^+	$(\text{Xe}_4^+)_{\text{III}} \text{Xe}_2$	3.644	4.40	0	-1.637	$\text{Xe}_4^+ \text{Xe} + \text{Xe}$	0.069
IX	Xe_6^+	$(\text{Xe}_3^+)_z (\text{Xe}_3)_{xz}$	3.375	4.20	0.021	-1.547	$\text{Xe}_3^+ \text{Xe}_2 + \text{Xe}$	0.108

Table 6. Transition energies T (eV) and moments μ (D) of the Xe_3^+Xe_k allowed transitions.

N ¹⁾	Xe_3^+Xe_k	Transitions					
		T	μ	T	μ	T	μ
II	Xe_3^+	1.87	11				
IV	Xe_3^+Xe	1.86	8.3	1.66	3.6	1.17	1.7
VII	Xe_3^+Xe_2	1.86	11			1.15	1.7
IX	Xe_3^+Xe_3	1.88	7.9	1.82	4.5	1.23	1.6

¹⁾ For N, see Table 6.

Table 7. Structure of the ionic clusters $\text{Xe}_k(\text{XeH})^+$ and ionic molecule $(\text{XeH})^+$.

$\text{Xe}_k(\text{XeH})^+$	$q_H^{1)}$	$R_{\text{Xe-H}}^{2)}$	$R_{\text{Xe-Xe}^+}^{2)}$	μ (D)	E	Xe Detachment
$(\text{XeH})^+$	0.282			1.54	-4.050	
$\text{Xe}(\text{XeH})^+$	0.280	6.07	4.46	11.5	-4.096	0.046
$\text{Xe}_2(\text{XeH})^+$	0.279	5.95	4.50	12.7	-4.156	0.060
$\text{Xe}_3(\text{XeH})^+$	0.278	5.96	4.56	13.3	-4.235	0.079
$\text{Xe}_4(\text{XeH})^+$	0.277	5.91	4.62	12.4	-4.297	0.062
$\text{Xe}_5(\text{XeH})^+$	0.275	5.73	4.85	10.7	-4.372	0.038

1) Charge of the H atom in $(\text{XeH})^+$.

2) $R_{\text{Xe-H}}$ and $R_{\text{Xe-Xe}^+}$ are the distances from the neutral atom to H and Xe of $(\text{XeH})^+$, respectively.

Figure Captions

- Fig. 1. Ar_3^+ and Ar_4^+ geometry. The Roman numerals indicate the clusters (the same as in Table 1). The numbers without a sign are distances in Å, and the numbers with the + sign are the atomic charges.
- Fig. 2. Geometry of the ionic clusters $\text{Ar}(\text{ArH})^+$, $\text{Ar}_2(\text{ArH})^+$ and ArH_2^+ . For notations, see Fig. 1.
- Fig. 3. Geometry of Xe_3^+ and Xe_4^+ . For notations, see Fig. 1. The Roman numerals indicate the clusters (see Table 5).
- Fig. 4. Geometry of the ionic clusters $\text{Xe}(\text{XeH})^+$, $\text{Xe}_2(\text{XeH})^+$, $(\text{XeH}_2)^+$ and Xe_2H_2^+ . For notations, see Figs. 1 and 3.

Fig. 1

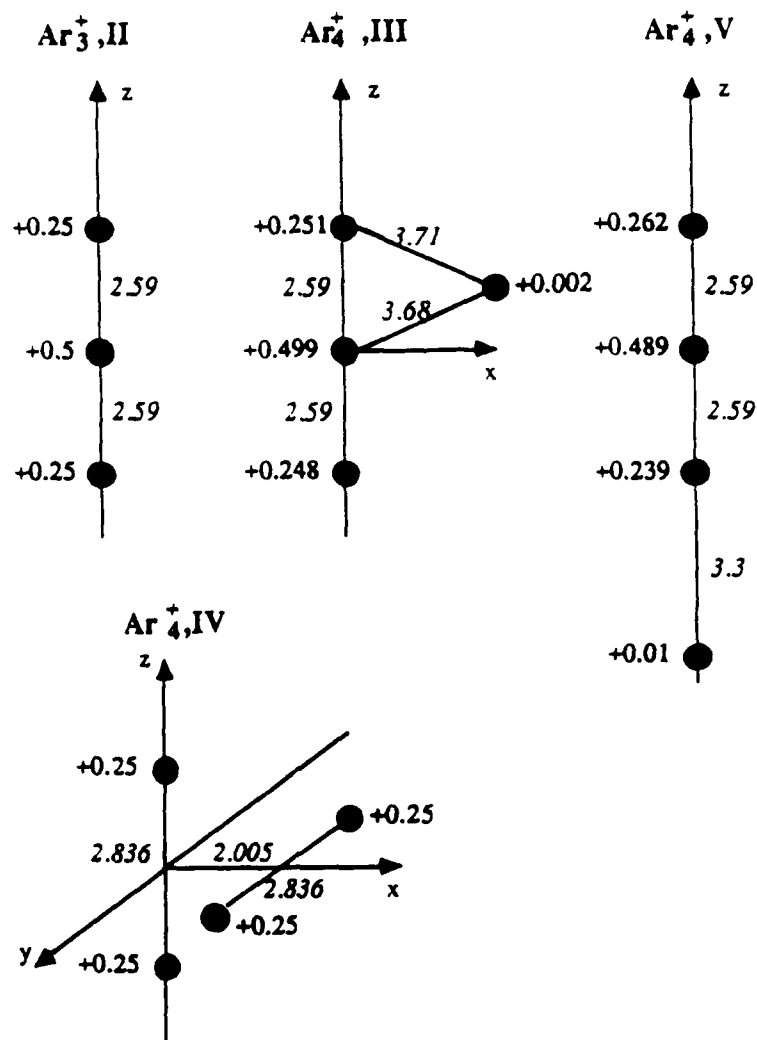


Fig. 2

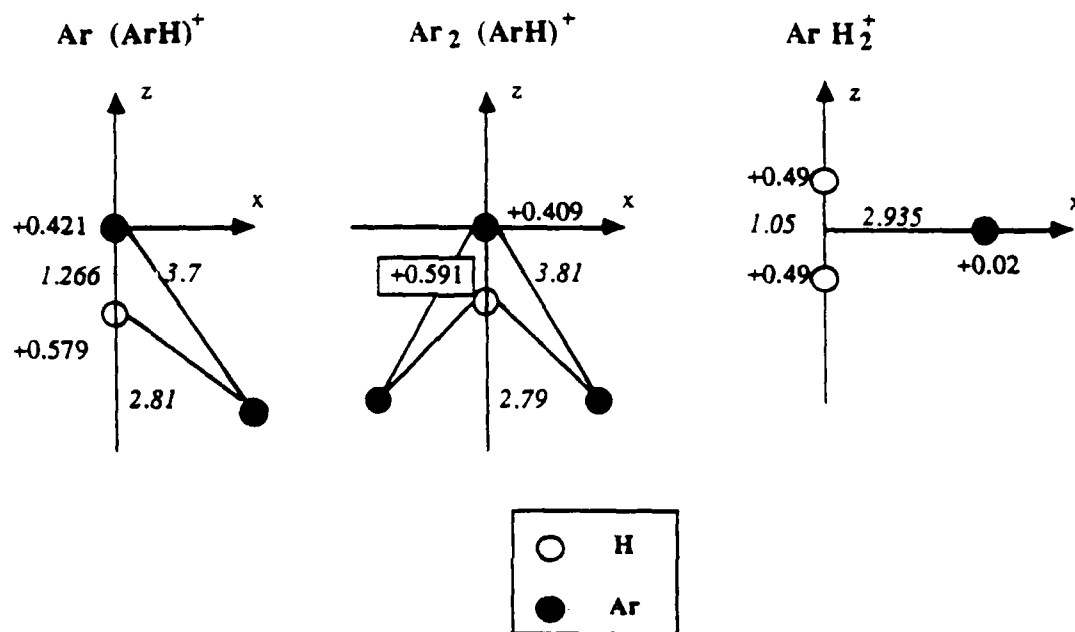


Fig. 3

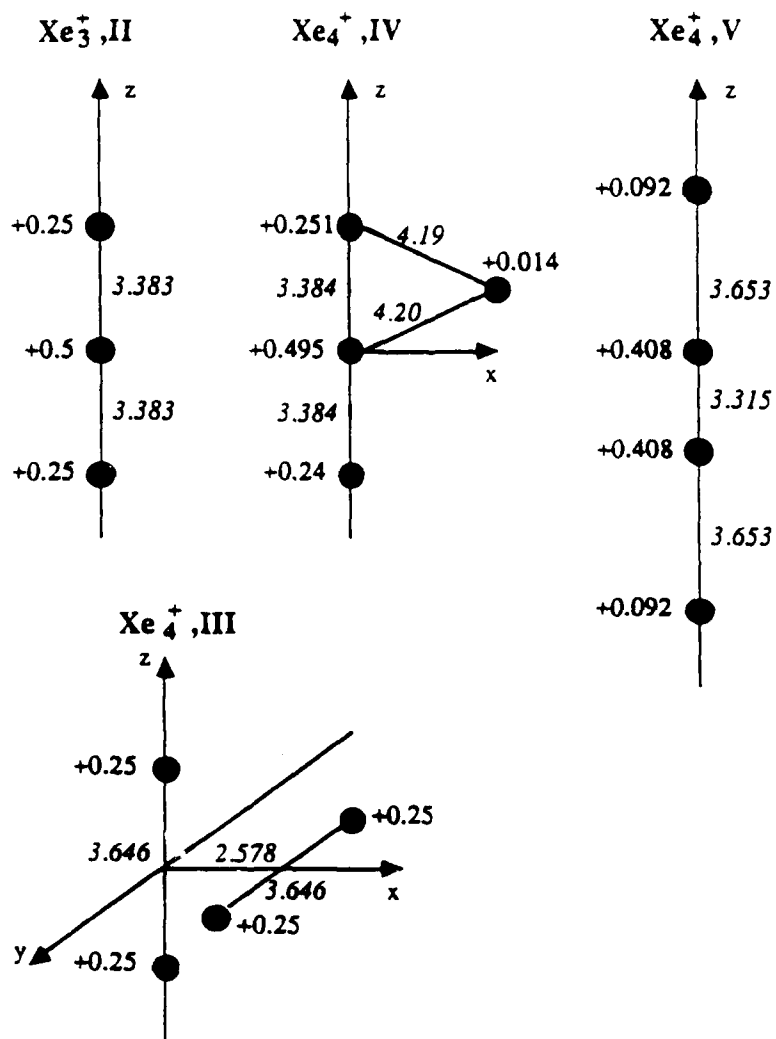
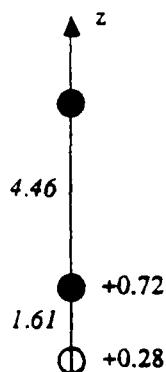
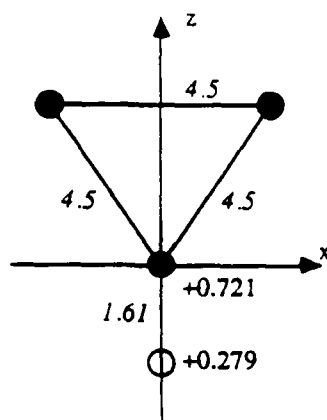


Fig. 4

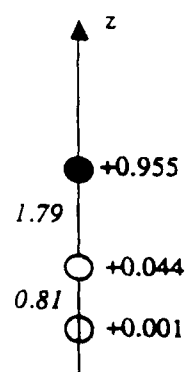
Xe (XeH)^+



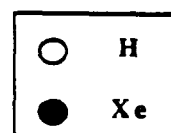
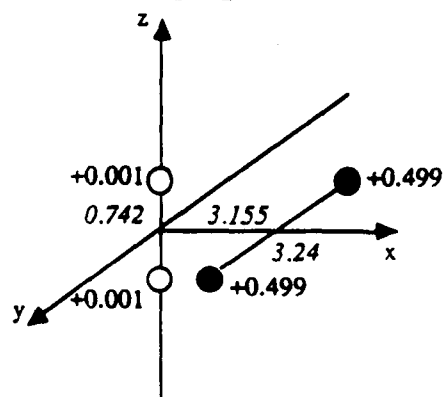
$\text{Xe}_2 (\text{XeH})^+$



$(\text{XeH}_2)^+$



$\text{Xe}_2^+ \text{H}_2$



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